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# Two-step Synthesis of New $\gamma$ -Lactones via Cyclization of 7-Chloro-2-(methoxycarbonyl)-4-6-dimethylocta-(2*E*,4*E*,6*E*)-trienoic acid

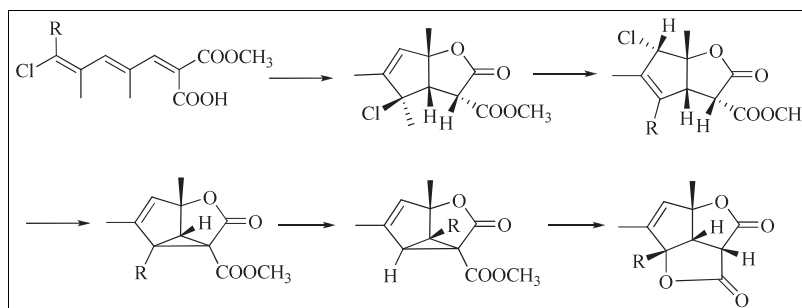
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A new rapid synthesis of  $\gamma$ -lactones, cis fused with a cyclopentenic ring by thermal cyclization of 7-chloro-2-(methoxycarbonyl)-4-6-dimethylocta-7-phenyl (or methyl) (2*E*,4*E*,6*E*)-trienoic acids was reported. The key step implicates an intramolecular cyclization to a cyclopentenyl cation, according to an electrocyclic  $\pi_{2s} + \pi_{2a}$  conrotatory process, published in a recent paper (from the corresponding diacids). We have investigated the thermal behavior of the corresponding half-esters since; if the cyclization obeys to the proposed mechanism, the diacids, half-esters must also cyclize in a similar manner. Saponification of these led to  $\gamma$ -dilactones via intermediary cyclopropanes. Mechanistic pathways were investigated.

## INTRODUCTION

In a recent paper, published in this journal [1], we report that thermal cyclization of 2-((2*E*,4*E*)-5-chloro-2,4-dimethylhexa-2,4-dienylidene)malonic acids led to new  $\gamma$ -dilactones (Scheme 1).

The key step implicates an intramolecular cyclization of the diacid to a cyclopentenyl cation, according to an electrocyclic  $\pi_{2s} + \pi_{2a}$  conrotatory process. The observed stereochemistry is that predicted by the Woodward–Hoffmann rules.

The scope of this new paper is to fix the limits of this reaction. So, in order to shed light on the suggested mechanistic pathways, we have investigated the thermal behavior of the corresponding half-esters. Since, if the cyclization obeys to the proposed mechanism in Scheme 2 [2], the diacids, half-esters must also cyclize in a similar manner. This new study permits us to isolate intermediates (unstable or stable), implied in the cyclization process.

Thus, thermal activation of **1a** and **1b** carried out in refluxing xylene (1 h, reflux) led to the previous  $\gamma$ -lactones-esters **2a** and **2b** (Scheme 2).

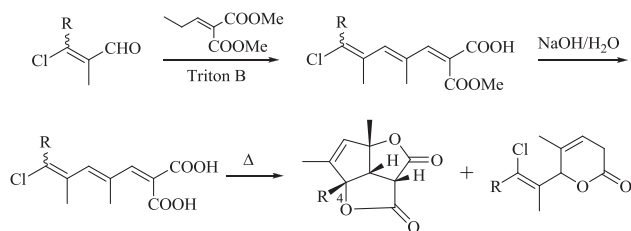
At lower temperature, in refluxing benzene, the less hindered compound **1b** (R=CH<sub>3</sub>) led in a closely quantitative yield, the  $\gamma$ -lactone **3b**, which could be considered as the kinetic product in the cyclization stage (Scheme 2). Hence, this latter rearranged in refluxing xylene (1 h, reflux) to the  $\gamma$ -lactone **2b** (Scheme 3).

The chlorolactone **3b** possessed effectively the stereochemistry predicted by the selection rules. Relative configuration of the bond C<sub>3a</sub>–C<sub>6a</sub> was set cis (for steric concerns), and the one at the bond C<sub>3</sub>–C<sub>3a</sub> was deduced due to its coupling constants ( $J = 8$  Hz), corresponding to a cis orientation [2,3]. C<sub>3a</sub>–C<sub>4</sub> configuration has been established by 1D Nuclear Overhauser Effect (NOE), which shown strong correlation of H<sub>3a</sub> and 6a-CH<sub>3</sub> and none between H<sub>3a</sub> and 4-CH<sub>3</sub>.

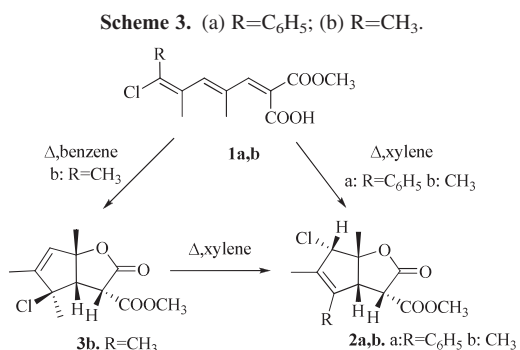
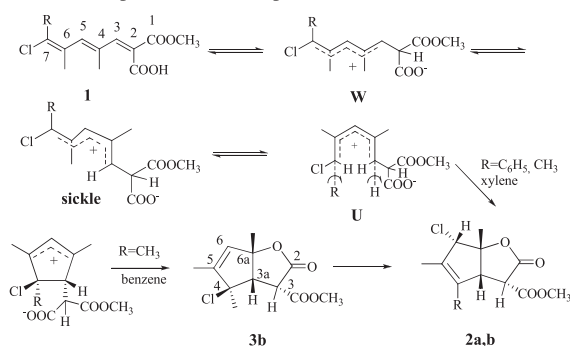
1,3 Transposition of the chlorine atom in lactone **3b** {**3b**  $\rightarrow$  **2b**} was thermally authorized by a *supra-antara* concerted mechanism, but, because of steric hindrances, only a few examples are known [4–6].

Dehydrohalogenation of lactones **2a** by anhydrous diethylamine afforded cyclopropanes **4a** (diethylamine 2 eq, room temperature (r.t.), 15 min). Compound **4a**, extremely unstable, rearranged slowly to its isomer **5a** (MeOH, reflux,

**Scheme 1.** **1** and **2**: a: R=C<sub>6</sub>H<sub>5</sub> (6E); b: R=C<sub>6</sub>H<sub>5</sub> (6Z); c: R=CH<sub>3</sub> (6E); d: R=CH<sub>3</sub> (6Z). **3**: R=C<sub>6</sub>H<sub>5</sub>; R=CH<sub>3</sub>.



**Scheme 2.** **1a** (6E): R=C<sub>6</sub>H<sub>5</sub>; **1a** (6Z): R=C<sub>6</sub>H<sub>5</sub>; **1b** (6E): R=CH<sub>3</sub>; **1b** (6Z): R=CH<sub>3</sub>; **2a**: R=C<sub>6</sub>H<sub>5</sub>; b: R=CH<sub>3</sub>; **3b**: R=CH<sub>3</sub>. Numbering of carbons as used for the comparison of NMR spectra.

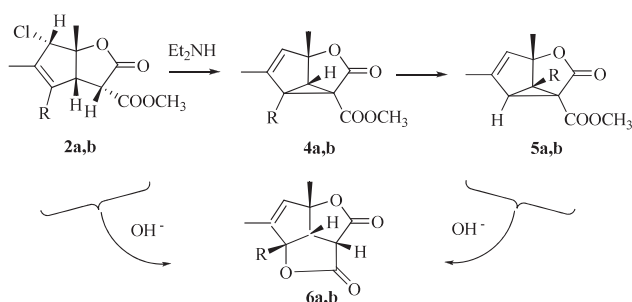


15 min). In the same conditions, **2b** gave a complex mixture, which slowly evolved to the dilactone **6b** (Scheme 4).

The 1,3-transposition of **4** → **5** seems to be particular to this series and represents, to our knowledge, the first example in the literature. Noteworthy, migration of a radical has been observed in cationic species [7,8]. Saponification (NaOH, CH<sub>3</sub>OH/H<sub>2</sub>O: 75/25, reflux 30 min) of lactones **2**–**5** afforded the dilactone **6** (Scheme 4).

Taking into account the results of the above figure, we may propose that formation of dilactones **6** in the saponification reaction could be processed via an intermediary cyclopropanic acid (not isolated) which rearrange spontaneously to the

**Scheme 4**



corresponding  $\gamma$ -lactone. These types of rearrangements have been reported (Scheme 5) [9,10].

In the same experimental conditions (and with various catalysis: paratoluenesulfonic acid or p-toluenesulfonic acid (PTSA), molecular sieves, bentonite, TiCl<sub>4</sub> etc.), the corresponding diester **7** (obtained by treating the sodium salts of diacids, half-esters by methyl iodide in HMPT) did not cyclize, implying, apart stereochemical implications, that at least one free carboxylic function was necessary (Scheme 6).

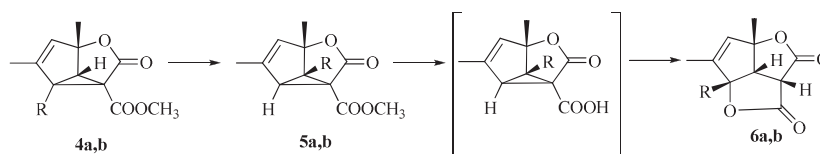
## EXPERIMENTAL

Melting points were taken on a Leitz 350 heated stage microscope (Ernst Leitz GMBH, Germany) and are not corrected. UV spectra were realized in ethanol, and  $\lambda_{\max}$  is given in nanometers ( $\epsilon$ ). NMR spectra (in CDCl<sub>3</sub>) were recorded on a Bruker Avance DPX 400 (Bruker Biospin S.A.S., France) and a WP80 DS instrument and were reported in parts per million (ppm) downfield from internal tetramethylsilane. Elemental analyses were indicated by elemental symbols. Dimethyl propylidene malonate and  $\beta$ -chloro  $\alpha$ -ethylenic aldehydes were synthesized to well-known procedures [11–17].

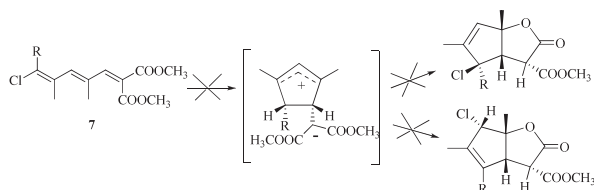
Characteristics of compounds **1** and **7** are in accordance with those described in the literature for related products [1]. Dimethyl propylidene malonate and  $\beta$ -chloro  $\alpha$ -ethylenic aldehydes were synthesized according to well-known procedures [1]. Malonic acid, half-esters **1**. About 150 mL of a methanolic triton B solution (40% by weight) was added to a mixture of aldehyde (0.1 mole) and dimethyl propylidene malonate (0.1 mole). The resulting solution was left at r.t. for 48 h. Dilution with 100 mL of water, extraction with ether of the by-products, and addition of 20% HCl to the aqueous layer provided the crude half-ester, which was extracted with ether. Sodium salts of **1a** and **1c** could be isolated by precipitation into a saturated solution of NaHCO<sub>3</sub>, washed with ether, and dried under reduced pressure. **1b** and **1d** were precipitated by acidification of the remaining soluble fraction, washed with water, and dried over MgSO<sub>4</sub>.

**1a** (E C6-C7): yield 90%; yellow crystals; mp 80°C (AcOEt/petroleum ether; b: benzene);  $\nu_{\text{CO}}$  1730, 1710;  $\lambda$

Scheme 5



Scheme 6



max 250, 320 (unstable at rt). *Anal.* calcd. for  $C_{17}H_{17}ClO_4$ : C, 63.64; H, 5.35; Cl, 11.05. Found: C, 63.47; H, 5.47; Cl, 11.03.  $^1H$  NMR: R s: 7.30; 6-CH<sub>3</sub> s: 2.15; 4-CH<sub>3</sub>: d 6.41 (1.75); H<sub>5</sub> m: 6.41; H<sub>3</sub> d: 7.20 (1); OCH<sub>3</sub> s: 3.75. **1b** (*Z* C6-C7): yield 40%; yellow crystals; mp 150°C (AcOEt/petroleum ether; b: benzene);  $\nu_{CO}$  1740, 1695;  $\lambda$  max 295 ( $\epsilon$ : 15,600). *Anal.* calcd. for  $C_{17}H_{17}ClO_4$ : C, 63.64; H, 5.35; Cl, 11.05. Found: C, 63.51; H, 5.48; Cl, 11.00.  $^1H$  NMR: R s: 7.40; 6-CH<sub>3</sub> s: 1.86; 4-CH<sub>3</sub>: d 1.91 (1); H<sub>5</sub> m: 6.73; H<sub>3</sub> d: 7.53 (1); OCH<sub>3</sub> s: 3.86. **1c** (*E* C6-C7): white crystals; mp 92°C (cyclohexane);  $\nu_{CO}$  1740, 1695;  $\lambda$  max 292 ( $\epsilon$ : 17,000). *Anal.* calcd. for  $C_{12}H_{15}ClO_4$ : C, 55.70; H, 5.85; Cl, 13.70. Found: C, 55.54; H, 5.96; Cl, 13.68.  $^1H$  NMR: R q: 1.90 (1.5); 6-CH<sub>3</sub> q: 2.03 (1.5); 4-CH<sub>3</sub>: d 1.83 (0.8); H<sub>5</sub> m: 6.40; H<sub>3</sub> d: 7.38 (0.8); OCH<sub>3</sub> s: 3.84. **1d** (*Z* C6-C7): yield 10%; White crystals; mp 114°C (cyclohexane);  $\nu_{CO}$  1720, 1695;  $\lambda$  max 287 ( $\epsilon$ : 14,000). *Anal.* calcd. for  $C_{12}H_{17}ClO_4$ : C, 55.70; H, 5.85; Cl, 13.70. Found: C, 55.59; H, 5.98; Cl, 13.58.  $^1H$  NMR: R q: 1.91 (1.5); 6-CH<sub>3</sub> q: 2.05 (1.5); 4-CH<sub>3</sub>: d 1.76 (0.8); H<sub>5</sub> m: 6.44; H<sub>3</sub> d: 7.43 (0.8); OCH<sub>3</sub> s: 3.86.

**$\gamma$ -Lactones 2a and 2b.** A solution of malonic acid, half-ester **1a** and **1b** (0.01 mole) in benzene (50 mL) was refluxed for 1 h. Distillation of the solvent under reduced pressure gave quantitatively the crude lactones-esters **2a** and **2b**.

**$\gamma$ -Lactone 3b.** A solution of malonic acid, half-ester **1c** (0.01 mole) in hexane (50 mL) was refluxed for 1 h. Distillation of the solvent under reduced pressure gave quantitatively the crude lactone-ester **3b**. **3b**: yield 100%; white crystals; mp 86°C (petroleum ether);  $\nu_{CO}$  1770, 1740;  $\lambda$  max 205 ( $\epsilon$ : 3500). *Anal.* calcd. for  $C_{12}H_{15}ClO_4$ : C, 55.70; H, 5.85; Cl, 13.70. Found: C, 55.59; H, 5.96; Cl, 13.57.  $^1H$  NMR: R s: 1.30; 5-CH<sub>3</sub> d: 1.38 (1.5); 6-CH<sub>3</sub> s: 1.47; H<sub>3</sub>: d 3.25 (8.5); H<sub>3a</sub> d: 3.66 (8.5); H<sub>6</sub> q: 5.13 (1.5); OCH<sub>3</sub> s: 3.38.  $^{13}C$  NMR 4-CH<sub>3</sub>: 25.4; 5-CH<sub>3</sub>: 11.6; 6a-CH<sub>3</sub>: 24.2; C<sub>3</sub>: 61.4; C<sub>3a</sub>: 53.4; C<sub>4</sub>: 77.8; C<sub>5</sub>: 146.4; C<sub>6</sub>: 131.4; C<sub>6a</sub>: 93.5; CO: 167.5; C<sub>3</sub>-COOCH<sub>3</sub>: 168.6; CH<sub>3</sub>O: 51.4.

**Rearrangement of lactone-ester 3b into lactone-ester 2b.** A solution of lactone-ester **3b** (0.01 mole) in xylene (50 mL) was refluxed for 1 h. Distillation of the solvent under reduced pressure gave practically quantitatively the crude lactone-ester **2b**, which recrystallized from petroleum ether. **2b**: yield 100%; white crystals; mp 118°C (petroleum ether);  $\nu_{CO}$  1770, 1740;  $\lambda$  max 205 ( $\epsilon$ : 3500). *Anal.* calcd. for  $C_{12}H_{15}ClO_4$ : C, 55.70; H, 5.85; Cl, 13.70. Found: C, 55.57; H, 5.97; Cl, 13.55.  $^1H$  NMR: R and 5-CH<sub>3</sub> m: 10.05 and 1.38; 6-CH<sub>3</sub> s: 1.65; H<sub>3</sub>: d 3.25 (2.3); H<sub>3a</sub> m: 3.13; H<sub>6</sub> m: 4.57 (1.5); OCH<sub>3</sub> s: 3.27.  $^{13}C$  NMR 4-CH<sub>3</sub> and 5-CH<sub>3</sub>: 12.0 and 12.3; 6a-CH<sub>3</sub>: 23.0; C<sub>3</sub>: 57.9; C<sub>3a</sub>: 50.7; C<sub>4</sub>: 133.4; C<sub>5</sub>: 134.8; C<sub>6</sub>: 73.8.4; C<sub>6a</sub>: 92.9; CO: 167.9; C<sub>3</sub>-COOCH<sub>3</sub>: 169.9; CH<sub>3</sub>O: 53.3.

**$\gamma$ -Lactone 2a.** Lactone **3a** (0.05 mole) in ether was treated by 2 eq of anhydrous diethylamine for 15 min at rt. The solution was then cooled, and the unstable lactone **4a** slowly crystallizes. **4a**: yield 80%; m.p.: 114°C;  $\nu_{CO}$  1770, 1720;  $^1H$  NMR: 4-CH<sub>3</sub> m: 7.33; 5-CH<sub>3</sub>: d: 1.63 (2); H<sub>6</sub> m: 5.63; 6a-CH<sub>3</sub> s: 1.75; H<sub>3a</sub> d: 3.74 (0.8); OCH<sub>3</sub> s: 3.44.  $^{13}C$  NMR: R (4-C<sub>6</sub>H<sub>5</sub>) 128.4, 129.0; 5-CH<sub>3</sub>: 15.2; 6a-CH<sub>3</sub> 19.9; C<sub>4</sub>: 55.7; C<sub>6</sub>: 132.6; C<sub>3</sub>-COOCH<sub>3</sub>: 169.5; OCH<sub>3</sub>: 52.7; CO: 167.2, 163.5.

**$\gamma$ -Lactone 5a.** The obtained lactone **4a** was heated in refluxing benzene for 30 min (or in MeOH, 60°C, 10 min). The benzene was distilled under reduce pressure, and the crude product was extracted with ether. After cooling, as mentioned earlier, lactone **5a** slowly crystallizes. **5a**: yield 80%; m.p.: 117°C;  $\nu_{CO}$  1770, 1730; *Anal.* calcd. for  $C_{17}H_{16}O_4$ : C, 71.82; H, 5.67; O, 22.51. Found: C, 71.64; H, 5.84; O, 22.70;  $^1H$  NMR R (C<sub>6</sub>H<sub>5</sub>) m: 7.33; 5-CH<sub>3</sub> m: 2.01; 6a-CH<sub>3</sub> s: 1.30; H<sub>4</sub> m: 5.55; H<sub>6</sub> m: 3.55; OCH<sub>3</sub> s: 3.58.  $^{13}C$  NMR R (C<sub>6</sub>H<sub>5</sub>) 128.6, 128.8, 129.27, 141.6; 5-CH<sub>3</sub>: 15.9; 6a-CH<sub>3</sub> 95.2; C<sub>4</sub>: 55.0; C<sub>6</sub>: 95.2; C<sub>3</sub>-COOCH<sub>3</sub>: 167.3; OCH<sub>3</sub> 52.6; CO: 163.6.

**$\gamma$ -Dilactones 6a and 6b.**  $\gamma$ -Lactones **3a**, **3b** and **5a**, **5b** (0.01 mole) were refluxed for 2 h in a solution of 2 M solution of sodium hydroxide (3 eq) in methanol/water (50/50). After acidification by a cold solution of 1-M HCl, lactones **6a** and **6b** were extracted with ether (3×50 mL). The solution was washed with water and dried over MgSO<sub>4</sub>. After filtration of the MgSO<sub>4</sub> and distillation of an half of ether, the  $\gamma$ -dilactones slowly crystallized from ether. **3a**: yield 90%; mp: 161°C (ether);  $\nu_{CO}$  1795, 1755;  $\lambda$  max 212 ( $\epsilon$ : 10.150). *Anal.* calcd. for  $C_{16}H_{14}O_4$ : C, 71.10; H, 5.22; O, 23.68.

Found: C, 70.89; H, 5.41; O, 23.78.  $^1\text{H}$  NMR R m: 7.40; 5-CH<sub>3</sub> d: 1.63 (1.3); 6a-CH<sub>3</sub> s: 1.77; H<sub>2a</sub> d: 4.05 (11.5); H<sub>3a</sub> q: 5.93 (1.3); H<sub>3</sub> d: 3.50 (11.5).  $^{13}\text{C}$  NMR **3a** R: 124.2; 128.5; 129.0; 146.7. 5-CH<sub>3</sub>: 12.2. 6a-CH<sub>3</sub> 24.8. 3-CH: 57.7. 4-C: 94.1; 3a-CH: 49.9. 6-CH: 133.3. 5-C-CH<sub>3</sub>: 138.3. 6a-CH<sub>3</sub>: 24.8. 4-C: 94.1. 6a-C: 94.3. CO: 167.9, 168.1. **3b**: yield 90%; mp: 123°C (pentane);  $\nu_{\text{CO}}$  1785, 1755;  $\lambda_{\text{max}}$  212 ( $\epsilon$ : 2.600). *Anal.* calcd. for C<sub>11</sub>H<sub>12</sub>O<sub>4</sub>: C, 63.45; H, 5.81; O, 30.74. Found: C, 63.29; H, 5.99; O, 30.08.  $^1\text{H}$  NMR R s: 1.60; 5-CH<sub>3</sub> d: 1.81 (1.3); 6a-CH<sub>3</sub> s: 1.60; H<sub>2a</sub> d: 4.08 (11.5); H<sub>3a</sub> q: 5.66 (1.3); H<sub>3</sub> d: 3.28 (11.5). **3b** R: 23.1. 5-CH<sub>3</sub>: 11.7. 6a-CH<sub>3</sub> 25.0. 4-C: 93.5; 3-CH: 55.4. 3a-CH: 50.6. 6-CH: 132.2. 5-C-H<sub>3</sub>: 146.7. 6a-CH<sub>3</sub>: 24.8. 4-C: 93.5. 6a-C: 94.1; CO: 168.1, 168.3.

**Diester 7.** About 0.01 mole of the malonic acid, half-ester **1** (sodium salt) was dissolved into 10 mL of hexamethylphosphoric triamide (HMPT). Then 0.01 mole of ICH<sub>3</sub> was added, and the solution was left at rt overnight. After acidification (HCl 20%), the crude mixture was extracted with ether, and the traces of malonic acid, half-ester were removed by a saturated solution of HCO<sub>3</sub>Na. The neutral fraction of diesters was purified by column chromatography (SiO<sub>2</sub> Merck 60F 254—ethyl acetate/cyclohexane—30:70). **7a** (*E*) yield: 100% (oil);  $\nu_{\text{CO}}$  1730, 1710;  $\lambda_{\text{max}}$  324 ( $\epsilon$ : 14,000), 254 (12,550). *Anal.* calcd. for C<sub>18</sub>H<sub>19</sub>ClO<sub>4</sub>: C, 64.56; H, 5.73; Cl, 10.58. Found: C, 64.34; H, 5.91; Cl, 10.58.  $^1\text{H}$  NMR R s: 7.28; 4-CH<sub>3</sub> s: 2.01 (1.3); 6a-CH<sub>3</sub> d: 1.61 (1.5); H<sub>5</sub> m: 6.36; H<sub>3</sub> d: 7.01 (1.5); OH<sub>3</sub> 2 s: ep 3.71. **7a** (*Z*) yield: 100% (oil);  $\nu_{\text{CO}}$  1720, 1605;  $\lambda_{\text{max}}$  290 ( $\epsilon$ : 13,300), 262 (11,300). *Anal.* calcd. for C<sub>18</sub>H<sub>19</sub>ClO<sub>4</sub>: C, 64.56; H, 5.73; Cl, 10.58. Found: C, 64.36; H, 5.88; Cl,

10.41.  $^1\text{H}$  NMR R s: 7.28; 4-CH<sub>3</sub> s: 1.86; 6a-CH<sub>3</sub> d: 1.91 (1.5); H<sub>5</sub> m: 6.66; H<sub>3</sub> d: 7.38 (1.5); OH<sub>3</sub> 2 s: 3.11, 3.86. **7b** (*E*) yield: 100% (oil);  $\nu_{\text{CO}}$  1720, 1610;  $\lambda_{\text{max}}$  290 ( $\epsilon$ : 15,700), 262 (11,300). *Anal.* calcd. for C<sub>13</sub>H<sub>17</sub>ClO<sub>4</sub>: C, 57.24; H, 6.29; Cl, 12.99. Found: C, 57.11; H, 6.43; Cl, 12.91.  $^1\text{H}$  NMR R q: 1.92 (1.5); 4-CH<sub>3</sub> q: 2.03 (1.5); 6a-CH<sub>3</sub> d: 1.75 (1.5); H<sub>5</sub> m: 6.41; H<sub>3</sub> d: 7.34 (1.5); OCH<sub>3</sub> 2 s: 3.91, 3.86.

## REFERENCES AND NOTES

- [1] Dillenschneider, J.-M.; Dufossé, L.; Fouillaud, M.; Sutthiwong, N.; Labia, R.; Valla, A. *J Heterocyclic Chem* 2015, in press.
- [2] Jackmann, L. M.; Sternell, S. *Applications of Nuclear Magnetic Resonance in Organic Chemistry*; Pergamon Press: London, 1972; Vol 5, 2nd ed., Chapter 4–2.
- [3] Anteunis, M.; Danneels, D. *Org Magn Reson* 1975, 7, 345.
- [4] Sprangler, C. W. *Chem Rev* 1976, 76, 187.
- [5] Strohmeier, W.; Eden, E. *Naturforsch* 1974, 29, 280.
- [6] Strohmeier, W. *Naturforsch* 1974, 29, 282.
- [7] Deno, N. C.; Pittman, C. U. *J Am Chem Soc* 1964, 86, 1871.
- [8] Deno, N. C.; Pittman, C. U.; Turner, J. O. *J Am Chem Soc* 1965, 87, 2153.
- [9] Den Besten, I. E.; Corkins, H. G. *J Chem Eng Data* 1970, 15, 453.
- [10] Bus, J.; Steinberg, H.; De Boer, T. J. *Rec Trav Pays Bas* 1972, 91, 657.
- [11] Corey, E. J. Ed. *Org Synth* 1966, 46, 18.
- [12] Arnold, Z.; Zemlicka, J. *Proc Chem Soc* 1958, 227.
- [13] Gagan, J. M.; Lane, A.; Lhord, O. *J Chem Soc* 1970, 2484.
- [14] Brettle, R.; Hydes, P. C.; Seddon, D.; Sutton, J. R.; Tooth, R. *J Chem Soc, Perkin Trans 1* 1973, 345.
- [15] Beauté, C.; Wolkowski, Z. W.; Merda, J. P.; Lelandais, D. *Tetrahedron Lett* 1971, 2473.
- [16] Wideqvist, S. *Ark Chem Miner Geol* 1945, B20, 1.
- [17] Greer, E. M.; Cosgriff, C. W. *Ann Rep Prog Chem, Sect B: Org Chem* 2013, 109, 328.